

Evolution of Water Chemistry from Surface Water to Aquifer

Logan Brink

NDSU

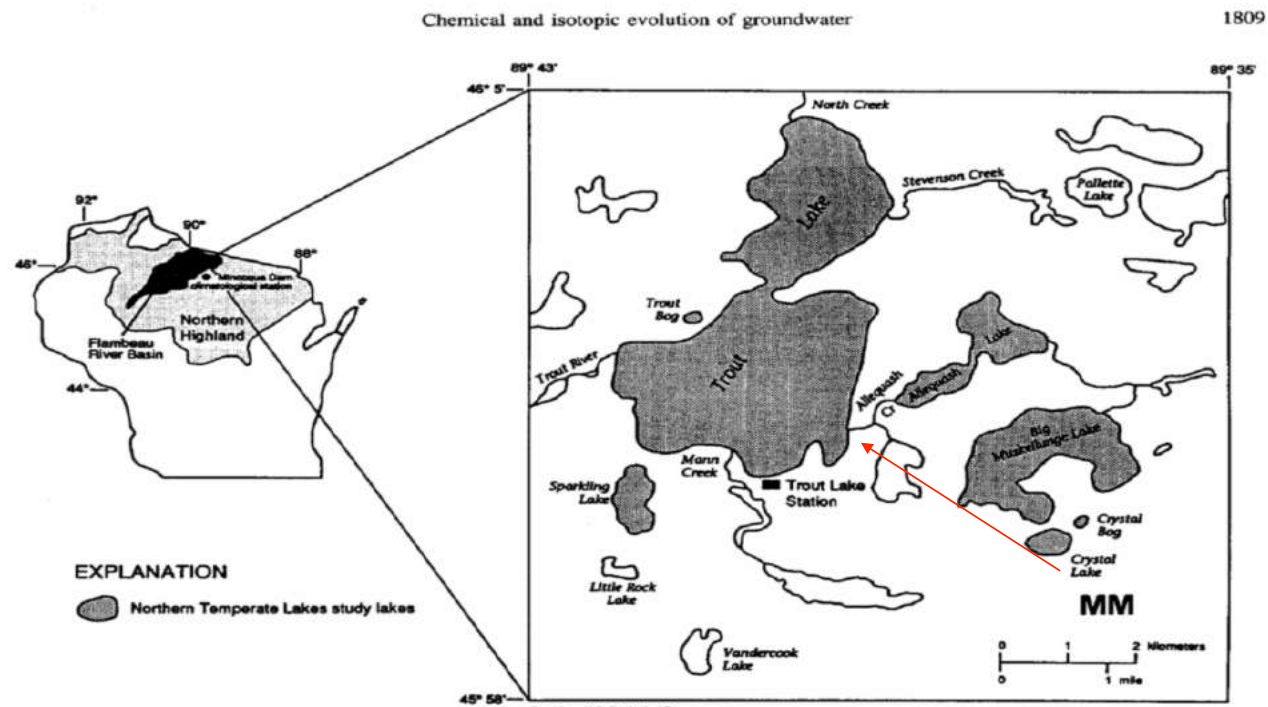
Geochemistry 428

12/06/2016

Background

- Two studies done
 - Galen J. Kenoyer, Carl J. Bowser
 - Thomas Bullen et al
- How chemistry of water changes as it flows to an aquifer
 - Kenoyer/Bowser one flow path
 - Bullen et al two separate paths

Trout Lake Watershed



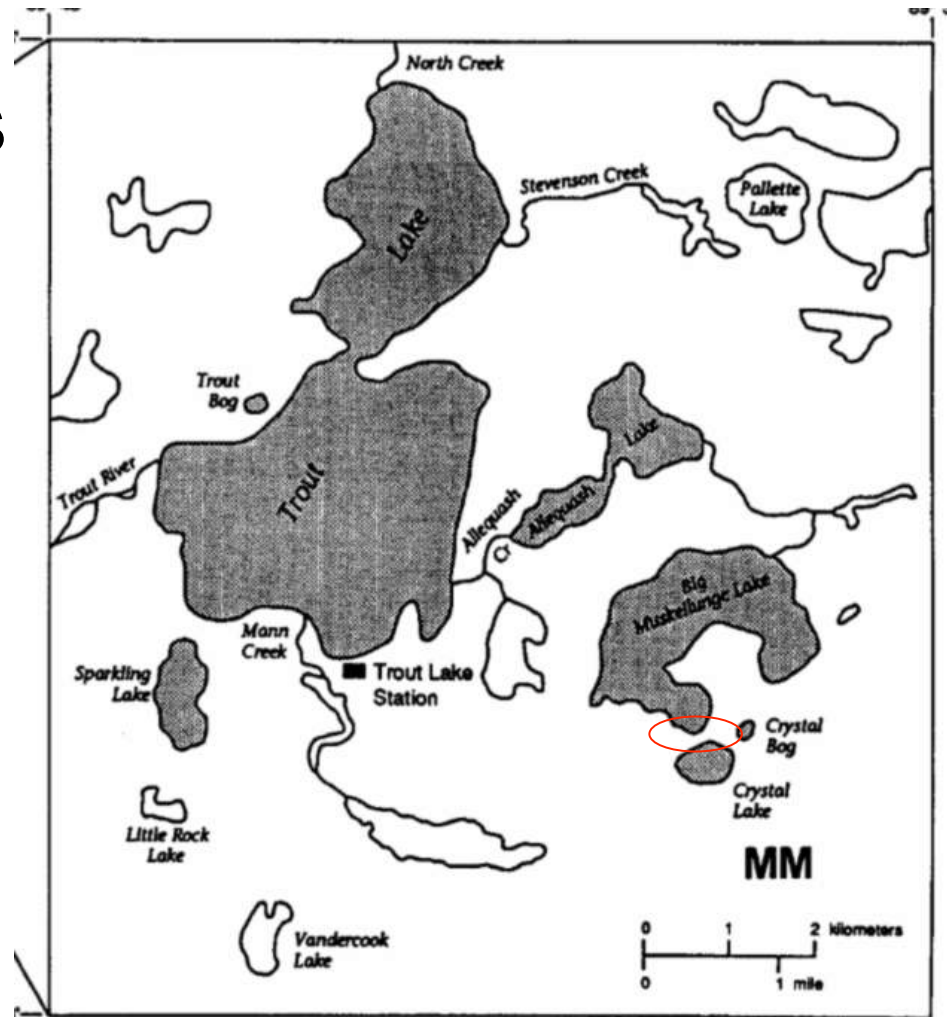
(Bullen et al, 1996)

Geology of Aquifer

- Aquifer- 50 m sandy glacial outwash
 - Copper Falls Formation
- Overlies Precambrian Rocks
 - Gneiss, amphibolite, schist, granite, monzonites, quartzose, and feldspathic sandstones.
- Groundwater is recharged by seepage from lakes, and by precipitation that infiltrates through highlands

Analysis of Isthmus

- Dominated by sand-sized materials
- Two extensive silt layers
- Minerals in sand fraction
 - 50-75% quartz
 - 9-15% calcic plagioclase
 - 9-15% sodic plagioclase
 - 2-10% feldspar
 - 1-2% biotite



- Clay minerals
 - 30-40% smectite
 - 30-40% illite
 - 15-30% chlorite/kaolinite
- Isthmus lies inside of state park
 - Compaction
 - Infiltration

Kenoyer-Bowser

- Flow from Crystal Lake to aquifer
- PCO_2 values range from 10^{-2} – $10^{-2.5}$
- Age impacts saturation of minerals
- Ion exchange is assumed to be constant

(Kenoyer, Bowser, 1992)

TABLE 1. Log (IAP/Ksp) for Selected Minerals and Waters Along a Groundwater Flow Path at Crystal Lake

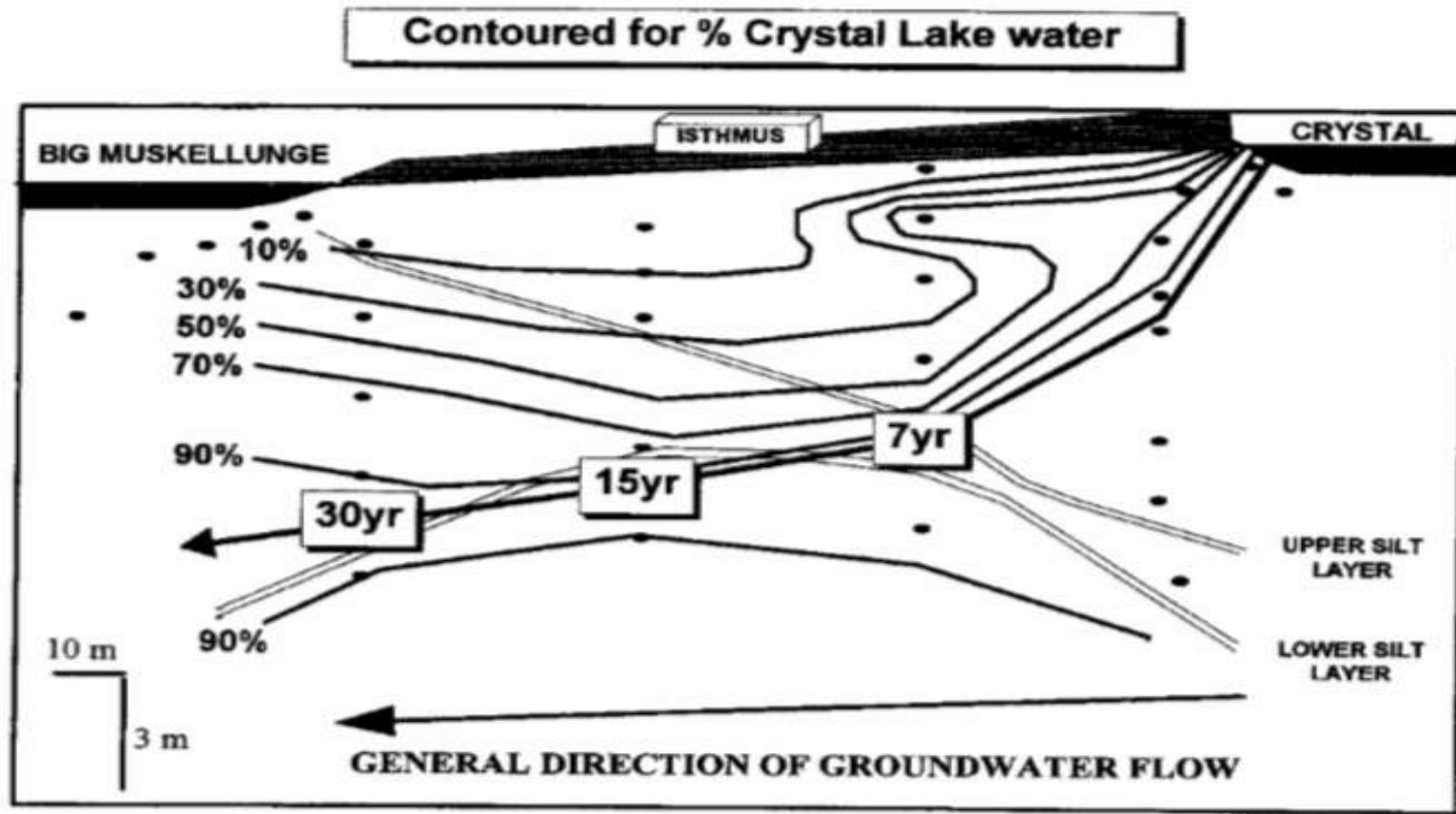
	Well Number		
	K86	K68	K75
Groundwater age, years	1.2	3.0	5.2
Albite	-7.5	-3.3	-3.1
Anorthite	-10.0	-5.0	-4.4
Phlogopite	-31	-25	-23
Diopside	-17	-14	-13
Forsterite	-21	-19	-18
Enstatite	-10	-8.6	7.8
Tremolite	-49	-39	-34
Calcite	-4.9	-4.3	-3.7
Dolomite	-10	-9.2	-8.0
Quartz	-0.3	0.4	0.4
Silica gel	-1.3	-0.6	-0.6
Halloysite	-2.5	1.9	1.9
Gibbsite	1	2	2
Kaolinite	2	6	6
Illite	-3	3	3
Ca-montmorillonite	-1	5	5
Montmorillonite (Belle Fourche)	-1	5	5
Montmorillonite (Aberdeen)	-2	3	4
Na-beidellite	-2	3	3

Kenoyer-Bowser

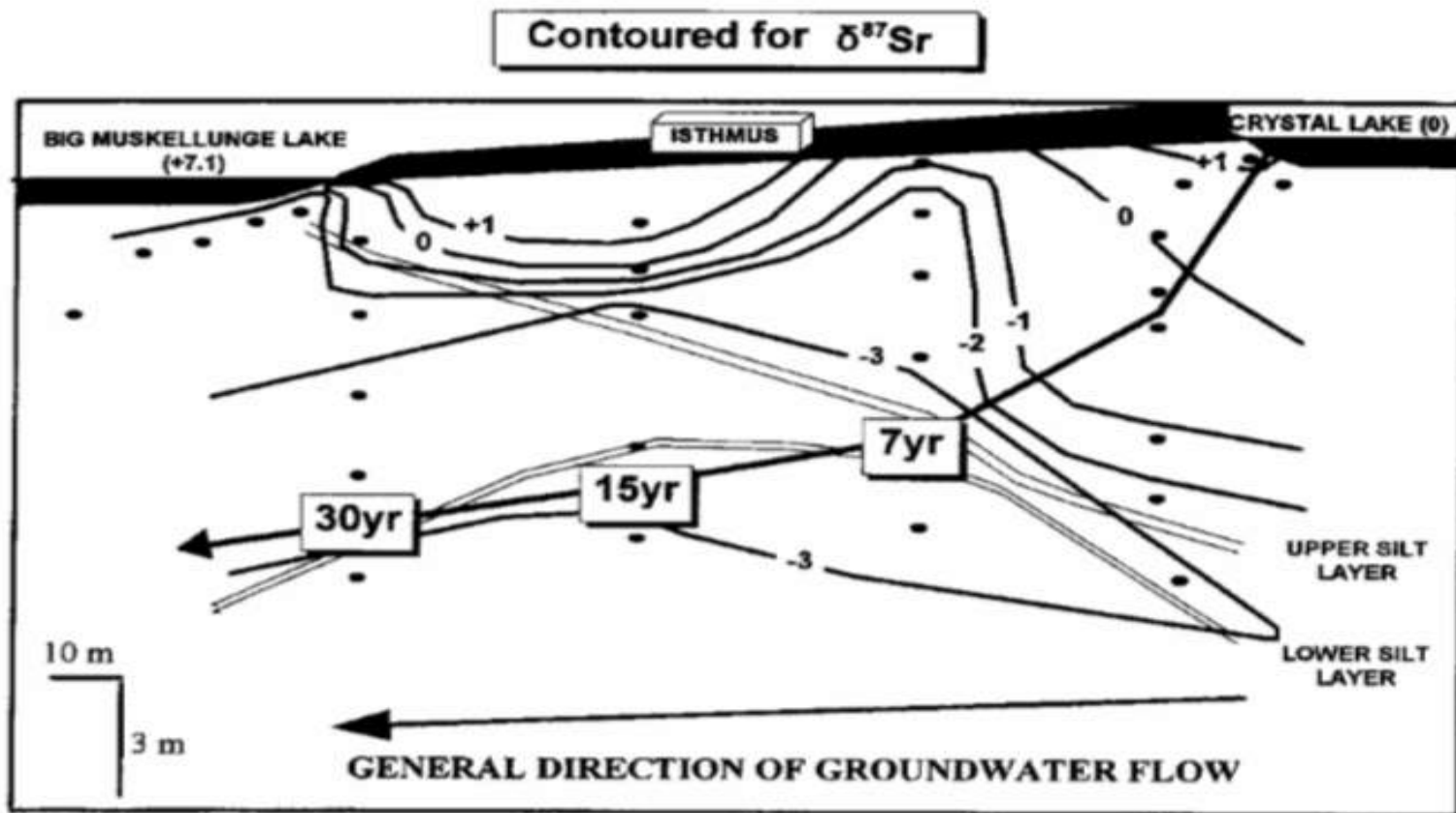
- Phreeqe
- Major increase in calcium, sodium, magnesium, silicon, and alkalinity
 - Explained by dissolution of plagioclase, biotite, and diopside
- Major chemical trends are explained by a model with those dissolutions and precipitation of kaolinite.
- Acidity of initial solution play a major role

Bullen's study

- Expanded on same idea
- Used Strontium isotopes to identify solute sources
 - Isotopes stable in solution and are persistent through chemical processes
 - Different minerals and age of sources vary isotope ratio
- Saturated lake pathway has high Sr concentration, compared to surface water on isthmus



(Bullen et al, 1996)



(Bullen et al, 1996)

Source of Sr

- Decrease in radiogenic Sr
 - Suggests Crystal lake path obtains Sr from Plagioclase and hornblende
- Upland pathways
 - Sr composition controlled by K-feldspar and biotite
 - Contain much more radiogenic Sr
- Different materials impact Sr differently, this allows for identification water source

TABLE 1. Chemical and strontium isotope data for isthmus groundwaters. Concentrations are in mg/L. Data are for the September, 1992 sampling except where labeled (6/93) which denotes samples collected during June, 1993. Samples are arranged from shallow to deep within individual nests.

sample	Ca	Mg	Sr	Na	K	Fe	Al	Si	⁸⁷ Sr/ ⁸⁶ Sr	δ ⁸⁷ Sr
nest 1:										
k1 (shallow)	4.12	0.92	0.017	1.98	0.51	3.00	0.01	0.87	0.71078	+0.76
k85	2.15	0.52	0.012	0.86	0.59	0.00	0.03	0.79	0.71024	0.00
k86	1.73	0.30	0.009	0.64	0.62	0.41	0.00	1.00	0.71018	-0.08
k86 (6/93)	1.96	0.33	0.009	0.57	0.53	0.47	0.04	0.92	0.71018	-0.08
k87	1.70	0.26	0.011	0.61	0.55	0.51	0.00	0.92	0.71008	-0.23
k87 (6/93)	1.89	0.37	0.011	0.59	0.52	0.69	0.23	1.25	0.71018	-0.08
k88	1.79	0.38	0.006	0.90	0.52	0.68	0.00	1.78	0.70949	-1.06
k89	2.81	0.48	0.018	0.95	0.67	0.12	0.00	2.90	0.70832	-2.70
k2 (deep)	5.37	1.97	0.017	1.66	0.59	0.03	0.01	5.08	0.70798	-3.18
nest 2:										
k66 (shallow)	9.42	1.61	0.068	1.53	2.61	26.94	0.02	3.68	0.70953	-1.00
k67	2.28	0.69	0.019	1.25	0.58	2.56	0.00	4.68	0.70812	-2.98
k68	4.26	0.99	0.027	1.35	0.74	0.00	0.00	4.26	0.70814	-2.96
k69	3.55	0.85	0.023	1.39	0.77	0.00	0.00	4.22	0.70815	-2.94
k70	2.62	0.70	0.017	1.20	0.54	0.21	0.00	4.08	0.70778	-3.46
k70 (6/93)	2.70	0.70	0.015	1.13	0.50	0.19	0.02	3.85	0.70786	-3.35
k71 (deep)	6.41	2.69	0.018	1.58	0.61	0.02	0.00	6.20	0.70741	-3.98
nest 3:										
k73 (shallow)	4.30	1.36	0.029	1.42	1.64	3.78	0.03	4.71	0.71213	+2.66
k74	3.56	1.36	nd	nd	nd	nd	nd	6.71	0.71026	+0.03
k75	2.39	0.65	0.018	1.49	0.58	4.75	0.00	5.17	0.70805	-3.08
k76	3.58	1.56	0.019	1.79	0.79	4.15	0.05	6.82	0.70771	-3.56
k76 (6/93)	3.51	1.50	0.018	1.89	0.64	5.91	0.01	7.23	0.70794	-3.24
k77 (deep)	3.09	1.03	0.006	0.40	1.60	0.58	0.11	0.34	0.71130	+1.49
k77 (6/93)	8.92	3.88	0.057	1.87	1.32	2.72	2.79	12.53	0.70827	-2.77
nest 4:										
k6 (shallow)	3.82	1.50	0.024	1.62	1.04	10.41	0.03	6.53	0.70969	-0.77
k80	3.72	1.93	0.012	2.56	0.56	6.15	0.02	7.55	0.70871	-2.15
k81	5.13	2.32	0.016	2.04	0.58	4.76	0.00	7.35	0.70782	-3.41
k82	7.05	3.25	0.018	1.77	0.58	1.94	0.01	11.03	0.70756	-3.77
k82 (6/93)	7.01	3.16	0.018	1.68	0.56	2.22	0.00	9.09	0.70756	-3.77
k5 (deep)	7.43	3.15	0.017	1.59	0.59	0.49	0.01	8.85	0.70825	-2.80
k5 (6/93)	7.94	3.23	0.017	1.43	0.63	0.42	0.00	8.41	0.70821	-2.86

(Bullen et al, 1996)

Bullen's conclusions

- Mineral dissolution reactions may change along flow paths, and differ along contrasting flow paths
- A broad range of radiogenic Sr was produced by weathering
- Compositions of Sr rely heavily on initial environmental conditions
- Contributions from different minerals change along a pathway, nature isn't a simple system

MFCRWSD



<http://websoilsurvey.sc.egov.usda.gov/App/WebSoilSurvey.aspx>



- Majority of area of interest is Korinis-Sunburg complex
 - Korinis soil accounts for 60%
 - Mollic Hapludalf
 - Glacial till

Map Unit Legend			
Kandiyohi County, Minnesota (MN067)			
Map Unit Symbol	Map Unit Name	Acres in AOI	Percent of AOI
L318A	Lundlake silty clay loam, 0 to 1 percent slopes	0.9	14.4%
L324A	Forestcity, overwash-Forestcity complex, 1 to 4 percent slopes	1.3	20.7%
L336A	Arctander, overwash-Arctander complex, 1 to 4 percent slopes	0.6	9.2%
L357C2	Korinis-Sunburg complex, 6 to 12 percent slopes, moderately eroded	3.4	55.2%
W	Water	0.0	0.5%
Totals for Area of Interest		6.2	100.0%

<http://websoilsurvey.sc.egov.usda.gov/App/WebSoilSurvey.aspx>

Assumptions

- Majority of infiltrating water is going to flow through Suburg-Koronis Complex
- Sundburg and Koronis soil units are similar in composition (Mollic Hapludalfs)
- Only taking into account infiltration of water from surface of isthmus
- New housing development doesn't impact infiltration or soil chemistry

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k87	1.70	0.26	0.011	0.61	0.55	0.51	0.00	0.92	0.71008	-0.23
k87 (6/93)	1.89	0.37	0.011	0.59	0.52	0.69	0.23	1.25	0.71018	-0.08
k88	1.79	0.38	0.006	0.90	0.52	0.68	0.00	1.78	0.70949	-1.06
k89	2.81	0.48	0.018	0.95	0.67	0.12	0.00	2.90	0.70832	-2.70
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k68	4.26	0.99	0.027	1.35	0.74	0.00	0.00	4.26	0.70814	-2.96
k69	3.55	0.85	0.023	1.39	0.77	0.00	0.00	4.22	0.70815	-2.94
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k82	7.05	3.25	0.018	1.77	0.58	1.94	0.01	11.03	0.70756	-3.77
k82 (6/93)	7.01	3.16	0.018	1.68	0.56	2.22	0.00	9.09	0.70756	-3.77
k5 (deep)	7.43	3.15	0.017	1.59	0.59	0.49	0.01	8.85	0.70825	-2.80
k5 (6/93)	7.94	3.23	0.017	1.43	0.63	0.42	0.00	8.41	0.70821	-2.86

Shallow sample

-----Saturation indices-----						
Phase	SI	log IAP	log KT			
Al(OH)3(a)	-8.25	2.55	10.80	Al(OH)3		
Albite	-15.60	-10.94	4.66	NaAlSi3O8		
Anorthite	-26.36	-0.74	25.62	CaAl2Si2O8		
Aragonite	-5.32	-13.66	-8.34	CaCO3		
Ca-Montmorillonite	-16.46	-8.67	7.78	Ca0.165Al2.33Si3.67O10(OH)2		
Calcite	-5.18	-13.66	-8.48	CaCO3		
CH4(g)	-123.23	-167.16	-43.93	CH4		
Chalcedony	-0.57	-4.12	-3.55	SiO2		
Chlorite(14A)	-67.20	1.18	68.38	Mg5Al2Si3O10(OH)8		
Chrysotile	-35.45	-3.25	32.20	Mg3Si2O5(OH)4		
CO2(g)	2.25	-15.90	-18.15	CO2		
Dolomite	-10.77	-27.86	-17.09	CaMg(CO3)2		
Fe(OH)3(a)	-0.36	17.55	17.91	Fe(OH)3		
Gibbsite	-5.56	2.55	8.11	Al(OH)3		
Goethite	5.58	17.60	12.02	FeOOH		
H2(g)	-37.84	-37.84	0.00	H2		
H2O(g)	-1.56	-0.05	1.51	H2O		
Hematite	13.21	35.24	22.03	Fe2O3		
Illite	-20.59	-8.72	11.86	K0.6Mg0.25Al2.35Si3.5O10(OH)2		
K-feldspar	-13.03	-10.94	2.09	KAlSi3O8		
K-mica	-18.44	-5.73	12.70	KAl3Si3O10(OH)2		
Kaolinite	-10.52	-3.08	7.43	Al2Si2O5(OH)4		
N2(g)	-0.66	-3.92	-3.26	N2		
NH3(g)	-54.20	-58.72	-4.52	NH3		
O2(g)	-7.54	75.58	83.12	O2		
Quartz	-0.14	-4.12	-3.98	SiO2		
Sepiolite	-24.89	-9.13	15.76	Mg2Si3O7.5OH:3H2O		
Sepiolite(d)	-27.79	-9.13	18.66	Mg2Si3O7.5OH:3H2O		
Siderite	-6.28	-17.17	-10.89	FeCO3		
SiO2(a)	-1.41	-4.12	-2.71	SiO2		
Strontianite	-6.87	-16.14	-9.27	SrCO3		
Talc	-32.83	-11.43	21.40	Mg3Si4O10(OH)2		

Deep Sample

-----Saturation indices-----						
Phase	SI	log IAP	log KT			
Aragonite	-5.72	-14.06	-8.34	CaCO3		
Calcite	-5.58	-14.06	-8.48	CaCO3		
CH4(g)	-119.51	-163.44	-43.93	CH4		
Chalcedony	-0.34	-3.89	-3.55	SiO2		
Chrysotile	-35.04	-2.84	32.20	Mg3Si2O5(OH)4		
CO2(g)	2.25	-15.90	-18.15	CO2		
Dolomite	-11.19	-28.28	-17.09	CaMg(CO3)2		
Fe(OH)3(a)	-3.72	14.19	17.91	Fe(OH)3		
Goethite	2.22	14.24	12.02	FeOOH		
H2(g)	-36.91	-36.91	0.00	H2		
H2O(g)	-1.56	-0.05	1.51	H2O		
Hematite	6.49	28.52	22.03	Fe2O3		
N2(g)	-0.66	-3.92	-3.26	N2		
NH3(g)	-52.80	-57.33	-4.52	NH3		
O2(g)	-9.39	73.73	83.12	O2		
Quartz	0.09	-3.89	-3.98	SiO2		
Sepiolite	-24.24	-8.48	15.76	Mg2Si3O7.5OH:3H2O		
Sepiolite(d)	-27.14	-8.48	18.66	Mg2Si3O7.5OH:3H2O		
Siderite	-9.18	-20.07	-10.89	FeCO3		
SiO2(a)	-1.18	-3.89	-2.71	SiO2		
Strontianite	-7.68	-16.95	-9.27	SrCO3		
Talc	-31.97	-10.57	21.40	Mg3Si4O10(OH)2		

Calculations

- Mollisols are high in organic material
 - C, P, N, K
- Same parent material
- Same physical properties
 - Majority sand sized particles
 - Some silt/clay
- Sr containing compound, less saturated in deeper sample

Conclusion

- Even though local soil is much higher in organic matter, some of the same concepts hold true.
 - Sr is less saturated as you go down in depth

References

- Bullen, T.D. Krabbenhoft, D.P. (1996). Kinetic and Mineralogic Controls on the Evolution of Groundwater Chemistry and $^{87}\text{Sr}/^{86}\text{Sr}$ in a Sandy Silicate Aquifer, Northern Wisconsin, USA. Pergamon, v. 60, pp. 1807-1821
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